12th Standard - Chemistry Practical

DATE

Simple salt analysis -I PRELIMINARY TESTS OBSERVATION

	DATE	PRELIMINARY TESTS	
S. No.	EXPERIMENT	OBSERVATION	INFERENCE
1	a)Colour Colour of the salt is noted	Colourless	Absence of copper and iron salt
	b)Appearance Appearance of the salt is noted	crystaline	May be Sulphate, Nitrate, Chloride
2	Solubility A little of the salt is shaken with water.	Soluble	May be Sulphate, Nitrate, Chloride orAmmonium Carbonate.
3	Action Of Heat: A small amount of the salt is heated gently in a dry test tube.	Decripitation occurs with evolution of reddish brown gas.	May be Nitrate .
4	Flame Test: A small amount of the salt is made into a paste withconc. HCl in a watch glass and introduced into the non-luminous part of the Bunsen flame.	No characteristic coloured flame	Absence of Calcium and Barium.
5	Ash Test: A filter paper is soaked into a paste of the salt withconc. HCl / HNO ₃ And Cobalt Nitrate solution in a watch glass and burnt.	No characteristic coloured ash.	Absence of Zinc, Aluminium and Magnesium.
		TICATION OF ACID RADICALS	
6	Action Of Dil Hcl To a small amount of dilute HCl the salt is added.	No characteristic change.	Absence of Sulphide and Carbonate.
7	Copper Turnings Test: A small amount of the salt is heated with Copper Turnings / Filter paper balland	Reddish brown gas is evolved.	Presence of Nitrate.
8	a few drops of conc. Sulphuric acid. Action Of NaOH: A small amount of the salt is heated with Sodium Hydroxide.	No pungent smelling gas.	Absence of Ammonium.
9	Chromyl Chloride Test: To a small amount of the salt a pinch of Potassium Dichromate is added and heated with few drops ofconc. Sulphuric acid.	No Red orange vapours.	Absence of chloride.
A sma	PARATION OF SODIUM CARBONATE EXTRACT: all amount of salt is mixed with twice the amount of tes, cooled and filtered. The filtrate is called "SOD!		illed water is added, boiled for 10
1	Barium Chloride Test: To a few drops of the extract,dilute Hydrochloric Acid is added until the effervescence ceases and 2 ml of Barium chloride solution is added.	No white precipitate	Absence of Sulphate
2	Silver Nitrate Test: To a few drops of the extract dilute Nitric Acid is added until the effervescence ceases and 2 ml of Silver Nitratesolution is added.	No precipitate.	Absence of chloride/ sulphide
3	Lead Acetate Test To a few drops of the extract dilute acetic acid is added until the effervesences ceases and 2 ml of Lead acetate solution is added	No precipitate.	Absence of sulphate and sulphide
1		D 1 1 C 1 .	Nituata is soutiums d
4	Brown Ring Test: To a few drops of extract dilute Sulphuric acid is added until the effervescence ceases, then freshly prepared FeSO ₄ is addedand then conctrated Sulphuric acid is added drop by drop along the sides of the test tube.	Brown ring is formed at the junction of the two layers.	Nitrate is confirmed.
PREI	To a few drops of extract dilute Sulphuric acid is added until the effervescence ceases, then freshly prepared FeSO ₄ is addedand then conctrated Sulphuric acid is added drop by drop along the sides of the test tube.	the junction of the two layers. ATION OF THE BASIC RADICALS	Nitrate is confirmed.
PREI	To a few drops of extract dilute Sulphuric acid is added until the effervescence ceases, then freshly prepared FeSO ₄ is addedand then conctrated Sulphuric acid is added drop by drop along the sides of the test tube. IDENTIFICATION OF ORIGINAL SOLUTION:	the junction of the two layers. ATION OF THE BASIC RADICALS	Absence of Ammonium

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	Hydroxide and Nessler's reagent are added.				
	GROUP SEPARATION				
2	To a few drops of the original solution 2 ml	white precipitate soluble when	Presence of First Group. (Lead).		
	of dilute HCl is added	boiled with water is obtained			
	CONFIRMAT	ORY TESTS FOR BASIC RADICALS			
1	To a few drops of the original	Yellow precipitatesoluble in hot	Lead is confirmed.		
	solution Potassium Iodide is added.	water which reappears as golden			
		yellow spangles on cooling.			

RESULT

The given simple salt contains

1. Basic Radical : Lead 2. Acid Radical : Nitrate

The given simple salt is : Lead Nitrate $\,$

Simple salt analysis -II

	DATE	PRELIMINARY TESTS	
S. No.	EXPERIMENT	OBSERVATION	INFERENCE
1	a)Colour Colour of the salt is noted	Colourless	Absence of copper and iron salt
	b)Appearance Appearance of the salt is noted	crystaline	May be Sulphate, Nitrate, Chloride
2	Solubility A little of the salt is shaken with water .	Soluble	May be Sulphate, Nitrate, Chloride o Ammonium Carbonate.
3	Action Of Heat: A small amount of the salt is heated gently in a dry test tube.	Salt sublimes with evolution of pungent smelling gas giving dense white fumes with a glass rod dipped in conc HCl.	May be Ammonium
4	Flame Test: A small amount of the salt is made into a paste with conc. HCl in a watch glass and introduced into the non-luminous part of the Bunsen flame.	No characteristic coloured flame	Absence of Calcium and Barium.
5	Ash Test: A filter paper is soaked into a paste of the salt with conc. HCl / HNO ₃ And Cobalt Nitrate solution in a watch glass and burnt.	No characteristic coloured ash.	Absence of Zinc, Aluminium and Magnesium.
	IDENTIF	ICATION OF ACID RADICALS	
6	Action Of Dil Hcl To a small amount of dilute HCl the salt is added.	No characteristic change.	Absence of Sulphide and Carbonate.
7	Copper Turnings Test: A small amount of the salt is heated with Copper Turnings and a few drops of conc. Sulphuric acid.	No reddish brown gas is evolved.	Absence of Nitrate.
8	Action Of NaOH: A small amount of the salt is heated with Sodium Hydroxide.	Pungent smelling gasforming dense whitefumes with a glass rod dipped in conc. HCl and also turns red Litmus paper blue.	Presence of Ammonium
9	Chromyl Chloride Test: To a small amount of the salt a pinch of Potassium Dichromate is added and heated with few drops of conc. Sulphuric acid.	No Red orange vapours.	Absence of chloride.
	RATION OF SODIUM CARBONATE EXTRACT:	1	1
	amount of salt is mixed with twice the amount		lled water is added, boiled for 10
minute 1	s, cooled and filtered. The filtrate is called "SOD Barium Chloride Test:		Sulphate is confirmed.
1	To a few drops of the extract, dilute Hydrochloric Acid is added until the	A white precipitate, insoluble in conc. HCl.	Suiphate is confirmed.
	effervescence ceases and 2 ml of Barium		

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	chloride solution is added.		
2	Silver Nitrate Test:	No precipitate.	Absence of chloride/ sulphide
	To a few drops of the extract dilute Nitric		
	Acid is added until the effervescence ceases		
	and 2 ml of Silver Nitrate solution is added.		
3	Lead Acetate Test	White precipitate soluble in	Presence of sulphate
	To a few drops of the extract dilute acetic	ammonium acetate and sodium	
	acid is added until the effervescence ceases	hydroxide is obtained	
	and 2 ml of Lead acetate solution is added		
4	Brown Ring Test:	No brown ring.	Absence of Nitrate.
	To a few drops of extract dilute Sulphuric		
	acid is added until the effervescence ceases,		
	then freshly prepared FeSO4 is added and		
	then conctrated Sulphuric acid is added		
	drop by drop along the sides of the test tube.		
	IDENTIFIC	ATION OF THE BASIC RADICALS	
	ARATION OF ORIGINAL SOLUTION:		
The or	rginal solution prepared by dissolving a smal	l amound of salt in water	
		ZERO GROUP	
1	To a few drops of the original	Reddish brown precipitate.	Ammonium is confirmed.
	solution Sodium Hydroxide and Nessler's		
	reagent are added.		
	. •	GROUP SEPARATION	
2	To a few drops of the original solution 2 ml	GROUP SEPARATION No characteristic precipitate.	Absence of First Group (Lead).
2			Absence of First Group (Lead).
2	To a few drops of the original solution 2 ml		Absence of First Group (Lead). Absence of Third Group
	To a few drops of the original solution 2 ml of dilute HCl is added	No characteristic precipitate.	
	To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml	No characteristic precipitate.	Absence of Third Group
3	To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OH solutions are added.	No characteristic precipitate. No characteristic precipitate.	Absence of Third Group (Aluminium).
3	To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OH solutions are added. To a few drops of the original solution 1 ml	No characteristic precipitate. No characteristic precipitate.	Absence of Third Group (Aluminium).
3	To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH ₄ Cl and 2 ml NH ₄ OH solutions are added. To a few drops of the original solution 1 ml NH ₄ Cl and 2 ml NH ₄ OH solutions are added	No characteristic precipitate. No characteristic precipitate.	Absence of Third Group (Aluminium).
3	To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH ₄ Cl and 2 ml NH ₄ OH solutions are added. To a few drops of the original solution 1 ml NH ₄ Cl and 2 ml NH ₄ OH solutions are added and H ₂ S gas is passed. To a few drops of the original solution 1 ml	No characteristic precipitate. No characteristic precipitate. No characteristic precipitate.	Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc).
3	To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH ₄ Cl and 2 ml NH ₄ OH solutions are added. To a few drops of the original solution 1 ml NH ₄ Cl and 2 ml NH ₄ OH solutions are added and H ₂ S gas is passed.	No characteristic precipitate. No characteristic precipitate. No characteristic precipitate.	Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc). Absence of Fifth Group
3	To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH ₄ Cl and 2 ml NH ₄ OH solutions are added. To a few drops of the original solution 1 ml NH ₄ Cl and 2 ml NH ₄ OH solutions are added and H ₂ S gas is passed. To a few drops of the original solution 1 ml NH ₄ Cl, 2 ml NH ₄ OH and 2 ml (NH ₄) ₂ CO ₃	No characteristic precipitate. No characteristic precipitate. No characteristic precipitate.	Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc). Absence of Fifth Group

CONFIRMATORY TESTS FOR BASIC RADICALS

Reddish brown precipitate.

Ammonium is confirmed.

RESULT

1

The given simple salt contains **1.** Basic Radical : **Ammonium**

solution Sodium Hydroxide and Nessler's

2. Acid Radical : Sulphate

Hydrogen Phosphate are added.

To a few drops of the original

reagent are added.

The given simple salt is: Ammonium Sulphate

Simple salt analysis -III

DATE PRELIMINARY TESTS S. No. **EXPERIMENT** OBSERVATION INFERENCE 1 a)Colour Colourless Absence of copper and iron salt Colour of the salt is noted b)Appearance Appearance of the salt is noted powdery May be carbonate and sulphide 2 **Solubility** insoluble May be carbonate and sulphide A little of the salt is shaken with water. 3 **Action Of Heat**: Colourless, odourlessgas turning May be Carbonate. A small amount of the salt is **heated** gently Lime water milky. in a dry test tube. 4 Flame Test: Brick red flame Presence of Calcium A small amount of the salt is made into a paste with conc. HCl in a watch glass and

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	introduced into the non-luminous part of the		
	Bunsen flame.		
5	Ash Test:	No characteristic coloured ash.	Absence of Zinc, Aluminium and
	A filter paper is soaked into a paste of the		Magnesium.
	salt with conc. HCl / HNO ₃		
	And Cobalt Nitrate solution in a watch glass and burnt.		
		ICATION OF ACID RADICALS	
6	Action Of Dil Hcl	Brisk effervescence of colourless,	Carbonate is confirmed.
	To a small amount of dilute HCl the salt is	odourless gas turning Lime	dar bonace is commined.
	added.	water milky.	
7	Copper Turnings Test:	No reddish brown gas is evolved.	Absence of Nitrate.
	A small amount of the salt is heated	_	
	with Copper Turnings and a few drops		
<u> </u>	of conc. Sulphuric acid.		
8	Action Of NaOH:	No pungent smelling gas.	Absence of Ammonium.
	A small amount of the salt is heated		
	with Sodium Hydroxide.		
9	Chromyl Chloride Test:	No Red orange vapours.	Absence of chloride.
	To a small amount of the salt a pinch of Potassium Dichromate is added and		
	heated with few drops of conc. Sulphuric		
	acid.		
PREPAI	RATION OF SODIUM CARBONATE EXTRACT:		
	amount of salt is mixed with twice the amount of	of sodium carbonate and 20ml of disti	lled water is added, boiled for 10
	s, cooled and filtered. The filtrate is called "SOD "		,
1	Barium Chloride Test:	No white precipitate	Absence of Sulphate
	To a few drops of the extract, dilute		
	Hydrochloric Acid is added until the		
	effervescence ceases and 2 ml of Barium		
	chloride solution is added.		
2	Silver Nitrate Test:	No precipitate.	Absence of chloride and sulphide
	To a few drops of the extract dilute Nitric		
	Acid is added until the effervescence ceases	1, 7,	
3	and 2 ml of Silver Nitrate solution is added.	No proginitate	Abanga of Culphata
3	Lead Acetate Test To a few drops of the extract dilute acetic	No precipitate	Absence of Sulphate
	acid is added until the effervescence ceases		
	and 2 ml of Lead acetate solution is added		
4	Brown Ring Test:	No brown ring.	Absence of Nitrate.
1	To a few drops of extract dilute Sulphuric	110010111111111111111111111111111111111	115561166 61 11161 6661
	acid is added until the effervescence ceases,		
	then freshly prepared FeSO4 is added and		
	then conctrated Sulphuric acid is added		
	drop by drop along the sides of the test tube.		
İ	IDENTIFIC	ATION OF THE DACK DADICALC	
		ATION OF THE BASIC RADICALS	
	RATION OF ORIGINAL SOLUTION:		
		amound of salt in dil HNO3	
The org	RATION OF ORIGINAL SOLUTION: ginal solution prepared by dissolving a small	amound of salt in dil HNO3 ZERO GROUP	Abcance of Ammonium
	RATION OF ORIGINAL SOLUTION: ginal solution prepared by dissolving a small To a few drops of the original	amound of salt in dil HNO3	Absence of Ammonium
The org	RATION OF ORIGINAL SOLUTION: ginal solution prepared by dissolving a small To a few drops of the original solution Sodium Hydroxide and Nessler's	amound of salt in dil HNO3 ZERO GROUP	Absence of Ammonium
The org	RATION OF ORIGINAL SOLUTION: ginal solution prepared by dissolving a small To a few drops of the original solution Sodium Hydroxide and Nessler's reagent are added.	amound of salt in dil HNO3 ZERO GROUP No Reddish brownprecipitate	Absence of Ammonium
The org	RATION OF ORIGINAL SOLUTION: ginal solution prepared by dissolving a small To a few drops of the original solution Sodium Hydroxide and Nessler's reagent are added.	amound of salt in dil HNO3 ZERO GROUP No Reddish brownprecipitate GROUP SEPARATION	
The org	RATION OF ORIGINAL SOLUTION: ginal solution prepared by dissolving a small To a few drops of the original solution Sodium Hydroxide and Nessler's reagent are added.	amound of salt in dil HNO3 ZERO GROUP No Reddish brownprecipitate	Absence of Ammonium Absence of First Group (Lead).
The org	To a few drops of the original solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution Sodium Hydroxide and Nessler's reagent are added.	amound of salt in dil HNO3 ZERO GROUP No Reddish brownprecipitate GROUP SEPARATION	
1 2	To a few drops of the original solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml of dilute HCl is added	amound of salt in dil HNO3 ZERO GROUP No Reddish brownprecipitate GROUP SEPARATION No characteristic precipitate.	Absence of First Group (Lead).
1 2	To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 2 ml of MILLE HCl and 2 ml NH4Cl and 2 ml NH4Cl of the original solution 3 ml of the original solution 1 ml NH4Cl and 2 ml NH4Cl of the original solution 1 ml	amound of salt in dil HNO3 ZERO GROUP No Reddish brownprecipitate GROUP SEPARATION No characteristic precipitate.	Absence of First Group (Lead). Absence of Third Group
1 2 3	To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added.	amound of salt in dil HNO3 ZERO GROUP No Reddish brownprecipitate GROUP SEPARATION No characteristic precipitate. No characteristic precipitate.	Absence of First Group (Lead). Absence of Third Group (Aluminium).
1 2 3 4	To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added.	amound of salt in dil HNO3 ZERO GROUP No Reddish brownprecipitate GROUP SEPARATION No characteristic precipitate. No characteristic precipitate. No characteristic precipitate.	Absence of First Group (Lead). Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc).
1 2 3	To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added and H2S gas is passed. To a few drops of the original solution 1 ml	amound of salt in dil HNO3 ZERO GROUP No Reddish brownprecipitate GROUP SEPARATION No characteristic precipitate. No characteristic precipitate.	Absence of First Group (Lead). Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc). Presence of Fifth
1 2 3 4	To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added and 2 ml NH4OHsolutions are added and 12 ml NH4OHsolutions are added and 12 ml NH4OH and 2 ml (NH4)2CO3	amound of salt in dil HNO3 ZERO GROUP No Reddish brownprecipitate GROUP SEPARATION No characteristic precipitate. No characteristic precipitate. No characteristic precipitate.	Absence of First Group (Lead). Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc).
1 2 3 4	To a few drops of the original solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added and H2S gas is passed. To a few drops of the original solution 1 ml NH4Cl, 2 ml NH4OH and 2 ml (NH4)2CO3 solutions are added.	amound of salt in dil HNO3 ZERO GROUP No Reddish brownprecipitate GROUP SEPARATION No characteristic precipitate. No characteristic precipitate. No characteristic precipitate. White precipitate	Absence of First Group (Lead). Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc). Presence of Fifth
1 2 3 4 5 5	To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added and H2S gas is passed. To a few drops of the original solution 1 ml NH4Cl, 2 ml NH4OH and 2 ml (NH4)2CO3 solutions are added.	amound of salt in dil HNO3 ZERO GROUP No Reddish brownprecipitate GROUP SEPARATION No characteristic precipitate. No characteristic precipitate. No characteristic precipitate. White precipitate ORY TESTS FOR BASIC RADICALS	Absence of First Group (Lead). Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc). Presence of Fifth Group (Calcium orBarium).
1 2 3 4	To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added and H2S gas is passed. To a few drops of the original solution 1 ml NH4Cl, 2 ml NH4OH and 2 ml (NH4)2CO3 solutions are added.	amound of salt in dil HNO3 ZERO GROUP No Reddish brownprecipitate GROUP SEPARATION No characteristic precipitate. No characteristic precipitate. No characteristic precipitate. White precipitate ORY TESTS FOR BASIC RADICALS White precipitate insoluble in	Absence of First Group (Lead). Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc). Presence of Fifth
1 2 3 4 5 5	To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added and H2S gas is passed. To a few drops of the original solution 1 ml NH4Cl, 2 ml NH4OH and 2 ml (NH4)2CO3 solutions are added.	amound of salt in dil HNO3 ZERO GROUP No Reddish brownprecipitate GROUP SEPARATION No characteristic precipitate. No characteristic precipitate. No characteristic precipitate. White precipitate ORY TESTS FOR BASIC RADICALS	Absence of First Group (Lead). Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc). Presence of Fifth Group (Calcium orBarium).

RESULT

The given simple salt contains

: Calcium 1. Basic Radical 2. Acid Radical : Carbonate
The given simple salt is : Calcium Carbonate

Simple salt analysis -IV PRELIMINARY TESTS

	DATE	Simple salt analysis –IV PRELIMINARY TESTS	
S. No.	EXPERIMENT	OBSERVATION	INFERENCE
0		020211111011	
1	a)Colour	Colourless	Absence of copper and iron salt
	Colour of the salt is noted		
	b)Appearance		
	Appearance of the salt is noted	crystaline	May be Sulphate, Nitrate, Chloride
2	Calabilita	C.1 hl.	M. L. C. L.L. A. NY.
2	Solubility A little of the salt is shaken with water.	Soluble	May be Sulphate, Nitrate, Chloride orAmmonium Carbonate.
3	Action Of Heat:	Decripitation occurs with	May be Nitrate .
3	A small amount of the salt is heated gently in	evolution of reddish brown gas.	May be Miliate.
	a dry test tube.	evolution of reduish brown gas.	
4	Flame Test:	No characteristic coloured flame	Absence of Calcium and Barium.
	A small amount of the salt is made into a		
	paste with conc. HCl in a watch glass and		
	introduced into the non-luminous part of the		
	Bunsen flame.		
5	Ash Test:	B <u>l</u> ue ash.	Presence of A <u>I</u> uminium.
	A filter paper is soaked into a paste of the		
	salt with conc. HCl / HNO ₃		
	And Cobalt Nitrate solution in a watch glass		
	and burnt.	ICATION OF ACID RADICALS	
6	Action Of Dil Hcl	No characteristic change.	Absence of Sulphide and
O	To a small amount of dilute HCl the salt is add		Carbonate.
7	Copper Turnings Test:	Reddish brown gas is evolved.	
,	A small amount of the salt is heated with Copp		Tresence on the deci
	Turnings / Filter paper balland a few drops		
	of conc. Sulphuric acid.		
8	Action Of NaOH:	No pungent smelling gas.	Absence of Ammonium.
	A small amount of the salt is heated with Sodi	um	
	Hydroxide.		
9	Chromyl Chloride Test:	No Red orange vapours.	Absence of chloride.
	To a small amount of the salt a pinch	٠.	
	of Potassium Dichromate is added and heate with few drops of conc. Sulphuric acid.	a	
DDEDA	RATION OF SODIUM CARBONATE EXTRACT:		
	amount of salt is mixed with twice the amount of	of sodium carbonate and 20ml of distil	led water is added hoiled for 10
	s, cooled and filtered. The filtrate is called "SODI		ica water is added, boiled for 10
1	Barium Chloride Test:	No white precipitate	Absence of Sulphate
	To a few drops of the extract, dilute	F F	F
	Hydrochloric Acid is added until the		
	effervescence ceases and 2 ml of Barium		
	chloride solution is added.		
2	Silver Nitrate Test:	No precipitate.	Absence of chloride/ sulphide
	To a few drops of the extract dilute Nitric Aci	d is	
	added until the effervescence ceases and 2 ml		
2	of Silver Nitratesolution is added. Lead Acetate Test	No proginitate	Absonge of sulphate and sulphate
3	To a few drops of the extract dilute acetic acid	No precipitate.	Absence of sulphate and sulphide
	added until the effervesences ceases and 2 ml	13	
	of Lead acetate solution is added		
4	Brown Ring Test:	Brown ring is formed at	Nitrate is confirmed.
-	To a few drops of extract dilute Sulphuric	the junction of the two layers.	1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1
	acid is added until the effervescence ceases,		
	then freshly prepared FeSO4 is addedand		
	then conctrated Sulphuric acid is added drop	by	
	drop along the sides of the test tube.		

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IDENTIFICATION OF THE BASIC RADICALS PREPARATION OF ORIGINAL SOLUTION: The orginal solution prepared by dissolving a small amound of salt in water ZERO GROUP To a few drops of the original solution **Sodium** No Reddish brown precipitate **Absence of Ammonium** 1 Hydroxide and Nessler's reagent are added. GROUP SEPARATION 2 To a few drops of the original solution 2 ml of No characteristic precipitate. Absence of First Group (Lead). dilute HCl is added 3 To a few drops of the original solution 1 ml **Gelatinous white**precipitate. Presence of **Third** NH₄Cl and 2 ml NH₄OH solutions are added. Group (Aluminium). **CONFIRMATORY TESTS FOR BASIC RADICALS** 1 To a few drops of the original A bright red lake. Aluminium is confirmed. solution Ammonium Hydroxide and Aluminon reagent are added.

RESULT

The given simple salt contains

1. Basic Radical : Aluminium

2. Acid Radical : Nitrate

The given simple salt is: **Aluminium Nitrate**

Simple salt analysis -V

DATE PRELIMINARY TESTS			
S. No.	EXPERIMENT	OBSERVATION	INFERENCE
1	a)Colour Colour of the salt is noted	Colourless	Absence of copper and iron salt
	b)Appearance Appearance of the salt is noted	crystaline	May be Sulphate, Nitrate, Chloride
2	Solubility A little of the salt is shaken with water.	Soluble	May be Sulphate, Nitrate, Chloride or Ammonium Carbonate.
3	Action Of Heat: A small amount of the salt is heated gently in a dry test tube.	Salt sublimes with evolution of pungent smelling gas giving dense white fumes with a glass rod dipped in conc HCl.	May be Ammonium
4	Flame Test: A small amount of the salt is made into a paste with conc. HCl in a watch glass and introduced into the non-luminous part of the Bunsen flame.	No characteristic coloured flame	Absence of Calcium and Barium.
5	Ash Test: A filter paper is soaked into a paste of the salt with conc. HCl / HNO ₃ And Cobalt Nitrate solution in a watch glass and burnt.	No characteristic coloured ash.	Absence of Zinc, Aluminium and Magnesium.
	IDENTIF	FICATION OF ACID RADICALS	
6	Action Of Dil Hcl To a small amount of dilute HCl the salt is added.	No characteristic change.	Absence of Sulphide and Carbonate.
7	Copper Turnings Test: A small amount of the salt is heated with Copper Turnings and a few drops of conc. Sulphuric acid.	No reddish brown gas is evolved.	Absence of Nitrate.
8	Action Of NaOH: A small amount of the salt is heated with Sodium Hydroxide.	Pungent smelling gasforming dense whitefumes with a glass rod dipped in conc. HCl and also turns red Litmus paper blue.	Presence of Ammonium
9	Chromyl Chloride Test:	Red orange vapoursevolved are	Chloride is confirmed.

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To a small amount of the salt a pinch of Potassium Dichromate is added and heated with few drops of conc. Sulphuric acid.

PREPARATION OF SODIUM CARBONATE EXTRACT:

passed through water to get a yellow solution, which on adding Lead acetateforms a yellowprecipitate

A small amount of salt is mixed with twice the amount of sodium carbonate and 20ml of distilled water is added, boiled for 10 minutes, cooled and filtered. The filtrate is called **"SODIUM CARBONATE EXTRACT"**.

IIIIIIuu	minutes, cooled and intered. The intrate is called SODIOM CARBONATE EXTRACT.				
1	Barium Chloride Test:	No white precipitate.	Absence of Sulphate		
	To a few drops of the extract, dilute				
	Hydrochloric Acid is added until the				
	effervescence ceases and 2 ml of Barium				
	chloride solution is added.				
2	Silver Nitrate Test:	A curdy white precipitate,	Presence of Chloride .		
	To a few drops of the extract dilute Nitric	soluble in excess of Ammonium			
	Acid is added until the effervescence ceases	hydroxide.			
	and 2 ml of Silver Nitrate solution is added.				
3	Lead Acetate Test	No precipitate	Absence of sulphate		
	To a few drops of the extract dilute acetic acid				
	is added until the effervescence ceases and 2				
	ml of Lead acetate solution is added				
4	Brown Ring Test:	No brown ring.	Absence of Nitrate.		
	To a few drops of extract dilute Sulphuric				
	acid is added until the effervescence ceases,				
	then freshly prepared FeSO 4 is added and				
	then conctrated Sulphuric acid is added				
	drop by drop along the sides of the test tube.				

IDENTIFICATION OF THE BASIC RADICALS

PREPARATION OF ORIGINAL SOLUTION:

The orginal solution prepared by dissolving a small amound of salt in water

		ZERO GROUP	
1	To a few drops of the original solution Sodium Hydroxide and Nessler's reagent are added.	Reddish brownprecipitate.	Ammonium is confirmed.
		GROUP SEPARATION	
2	To a few drops of the original solution 2 ml of dilute HCl is added	No characteristic precipitate.	Absence of First Group (Lead).
3	To a few drops of the original solution 1 ml NH₄Cl and 2 ml NH₄OH solutions are added.	No characteristic precipitate.	Absence of Third Group (Aluminium).
4	To a few drops of the original solution 1 ml NH ₄ Cl and 2 ml NH ₄ OH solutions are added and H ₂ S gas is passed.	No characteristic precipitate.	Absence of Fourth Group (Zinc).
5	To a few drops of the original solution 1 ml NH ₄ Cl, 2 ml NH ₄ OH and 2 ml (NH ₄) ₂ CO ₃ solutions are added.	No characteristic precipitate	Absence of Fifth Group (Calcium and Barium).
6	To a few drops of the original solution 1 ml NH ₄ Cl, 2 ml NH ₄ OH and 2 ml Di Sodium Hydrogen Phosphate are added.	No characteristic precipitate.	Absence of Sixth Group (Magnesium)
	CONFIRMAT	ORY TESTS FOR BASIC RADICALS	
1	To a few drops of the original solution Sodium Hydroxide and Nessler's reagent are added.	Reddish brown precipitate.	Ammonium is confirmed.

RESULT

The given simple salt contains

1. Basic Radical : Ammonium 2. Acid Radical : Chloride

The given simple salt is : $\mbox{\bf Ammonium Chloride}$

Simple salt analysis -VI

DATE PRELIMINARY TESTS

S. No.	EXPERIMENT	OBSERVATION	INFERENCE
1	a)Colour Colour of the salt is noted	Colourless	Absence of copper and iron salt

DiAppearance DiAppearance Appearance		<u>www.Padasalal.Net</u>		<u>www.rrbrnpsc.com</u>
A little of the salt is shaken with water. 3			crystaline	May be Sulphate, Nitrate, Chloride
Action Of Heat A small amount of the salt is heated gently in a dry test tube. Flame Test: A small amount of the salt is made into a paste with conc. HCI in a watch glass and introduced into the non-luminous part of the Binnsen flame. A sh Test: A filter paper is soaked into a paste of the salt with conc. HCI HNO3 And Cobah Nitrate solution in a watch glass and purit. DENTIFICATION OF ACID RADICALS A CLOON DIJ HCI To a small amount of dilute HCI the salt is added. A cloon Of Dij HCI To a small amount of the salt is heated with Sodium Hydroxide. B Action Of NaOB: A chinn Great salt with conc. HCI A copper unrings Test: A small amount of the salt is heated with Sodium Hydroxide. B Action Of NaOB: A small amount of the salt is heated with Sodium Hydroxide. B Action Of NaOB: A small amount of the salt a pinch of Potatssium Dictromate is added and heated with few drops of conc. Sulphuric acid. PREPARATION OF SODIUM CARBONATE EXTRACT: A small amount of the salt a pinch of the salt is indiced with twice the amount of sodium carbonate and 20ml of distilled water is added, builded for 10 minutes, cnoled and filtered. The filtrate is called "SODIUM CARBONATE EXTRACT". B Barium Chloride Test: To a few drops of the extract, dilute Hydrochioric Acid is added until the effervescence ceases and 2 ml of Silver Nitrate that illus a seid: acid monthly of the salt is mixed with twice the amount of the serves cance and 2 ml of Silver Nitrate that illus a seid: acid monthly of the serves and 2 ml of Silver Nitrate that illus a seid: acid monthly of the serves and 2 ml of Silver Nitrate that illus a seid: acid and filtered; acid is added and illusted with the discress cance and 2 ml of Barium chloride test in the serves and 2 ml of Barium chloride acid is added and illusted with the discress and 2 ml of Silver Nitrate (acid is added and the concretated subjustic acid is added and the concretat	2		Soluble	
A small amount of the salt is made into a paste with cone. HCI in a vacte plass and introduced into the non-luminous part of the Blunsen flame. 5. Ash Test: A filter paper is soaked into a paste of the salt with cone. HCI / HNO) And Cobalt Nitrate solution in a watch glass and burnt. 10ENTIFICATION OF ACID RADICALS 6. Action Of Dil Hcl To a small amount of dilute HCI the salt is added. 7. Copper Turnings Test: A small amount of the salt is heated with Copper Turnings and a few drops of cone. Sulphuric acid. A small amount of the salt is heated with Sodium Hydroxide. 9. Chromyl Chloride Test: To a small amount of the salt a pinch of Potassium Dichromate is added and heated with few drops of cone. Sulphuric acid. PREPARATION OF SODIUM CARBONATE EXTRACT: A small amount of the salt a pinch of Potassium Dichromate is added until the effervescence cases and 2 ml of Barium chloride Solution is added. 1. Barium Chloride Test: To a few drops of the extract, dilute Nitric Acid is added until the effervescence cases and 2 ml of Sarium chloride added until the effervescence cases and 2 ml of Sarium chloride Test: To a few drops of the extract dilute Nitric Acid is added until the effervescence cases and 2 ml of Sarium chloride Test: To a few drops of the extract dilute Nitric Acid is added until the effervescence cases and 2 ml of Sarium chloride Test: To a few drops of the extract dilute Nitric Acid is added until the effervescence cases and 2 ml of Sarium chloride Test: To a few drops of the extract dilute Nitric Acid is added until the effervescence ceases and 2 ml of Sarium chloride Test: To a few drops of the extract dilute Nitric Acid is added until the effervescence ceases and 2 ml of Sarium chloride Test: To a few drops of the extract dilute Sulphuric acid is added until the effervescence ceases and 2 ml of Barium chloride Test: To a few drops of the office the salt is added until the effervescence ceases and 2 ml of Sarium chloride Test. No brown ring. No brown ring. No brown ring. Absence of Nitrate.	3	Action Of Heat: A small amount of the salt is heated gently	of pungent smelling gas giving dense white fumes with a glass rod dipped	
Affilter paper is soaked into a paste of the salt with conc. HCJ PNO). And Cobalt Nitrate solution in a watch glass and burnt. IDENTIFICATION OF ACID RADICALS	4	A small amount of the salt is made into a paste with conc. HCl in a watch glass and introduced into the non-luminous part of the	No characteristic coloured flame	Absence of Calcium and Barium.
Brisk effervescence of colourless, added. Brisk effervescence of colourless, added. Assence of Nitrate.	5	A filter paper is soaked into a paste of the salt with conc. HCl / HNO ₃ And Cobalt Nitrate solution in a watch glass and burnt.		
To a small amount of dilute HCl the salt is added. 7 Copper Turnings Test: A small amount of the salt is heated with Copper Turnings and a few drops of conc. Sulphuric acid. 8 Action Of NaOH: A small amount of the salt is heated with Sodium Hydroxide. 9 Chromyl Chloride Test: To a small amount of the salt a pinch of Pressure of Conc. Sulphuric acid. 9 Chromyl Chloride Test: To a small amount of the salt a pinch of Pressure of Conc. Sulphuric acid. 9 Chromyl Chloride Test: To a small amount of the salt a pinch of Pressure of Conc. Sulphuric acid. 9 Chromyl Chloride Test: To a few drops of the extract, dilute Hydrochloric Acid is added and heated with few drops of conc. Sulphuric acid. 1 Barium Chloride Test: To a few drops of the extract, dilute Hydrochloric Acid is added until the effervescence ceases and 2 ml of Barium chloride solution is added. 3 Lead Acetate Test: To a few drops of the extract dilute acid is added until the effervescence ceases and 2 ml of Silver Nitrate solution is added. 4 Brown Ring Test: To a few drops of the extract dilute acid is added until the effervescence ceases and 2 ml of Silver Nitrate solution is added. 4 Brown Ring Test: To a few drops of the extract dilute acid is added until the effervescence ceases and 2 ml of Silver Nitrate solution is added. 4 Brown Ring Test: To a few drops of the extract dilute acid is added until the effervescence ceases, then freshly prepared FeSO is added and then conctrated Sulphuria caid is added drop by drop along the sides of the test tube. DENTIFICATION OF THE BASIC RADICALS PREPARATION OF ORIGINAL SOLUTION: The orginal solution prepared by dissolving a small amound of salt in water ZERO GROUP 1 To a few drops of the original solution Sodium Hydroxide and Nessler's regent are added. Ammonium is confirmed.		IDENTIF	ICATION OF ACID RADICALS	
Copper Turnings Test: A small amount of the salt is heated with Copper Turnings and a few drops of conc. Sulphuric acid. Pungent smelling gas/forming dense white/Immes with a glass rod dipped in conc. HCl and also turns red Litmus paper blue. Presence of Ammonium gas/forming dense white/Immes with a glass rod dipped in conc. HCl and also turns red Litmus paper blue. Absence of chloride. Presence of Chloride. Presence of Chloride. Presence of Chloride of Potassium Dichromate is added and heated with few drops of conc. Sulphuric acid. No Red orange vapours. Absence of chloride. Absence of chloride.	6	To a small amount of dilute HCl the salt is added.	odourless gas turning Lime	Carbonate is confirmed.
Action of NaOH: A small amount of the salt is heated with Sodium Hydroxide. Programming dense whitefumes with a glass rod dipped in conc. HCl and also turns red Litmus paper blue. Proceedings of Chromyl Chloride Test: To a small amount of the salt a pinch of Potassium Dichromate is added and heated with few drops of conc. Sulphuric acid. No Red orange vapours. Absence of chloride.	7	A small amount of the salt is heated with Copper Turnings and a few drops		Absence of Nitrate.
Silver Nitrate Test: To a few drops of the extract dilute Nitric Acid is added until the effervescence ceases and 2 ml of Silver Nitrate solution is added until the effervescence ceases and 2 ml of Lead acetate solution is added until the effervescence ceases and 2 ml of Lead acetate solution is added until the effervescence ceases then freshly prepared FeSol is added and then conctrated Sulphuric acid is added until the effervescence ceases then freshly prepared feSol is added and then conctrated Sulphuric acid is added. No brown ring. No brown ring. Absence of Nitrate.	8	Action Of NaOH: A small amount of the salt is heated	gasforming dense whitefumes with a glass rod dipped in conc. HCl and also turns red Litmus	Presence of Ammonium
A small amount of salt is mixed with twice the amount of sodium carbonate and 20ml of distilled water is added, boiled for 10 minutes, cooled and filtered. The filtrate is called "SODIUM CARBONATE EXTRACT". Barium Chloride Test:	9	To a small amount of the salt a pinch of Potassium Dichromate is added and heated with few drops of conc. Sulphuric	No Red orange vapours.	Absence of chloride.
Barium Chloride Test: To a few drops of the extract, dilute Hydrochloric Acid is added until the effervescence ceases and 2 ml of Barium chloride solution is added.	A small	l amount of salt is mixed with twice the amount o		lled water is added, boiled for 10
To a few drops of the extract, dilute Hydrochloric Acid is added until the effervescence ceases and 2 ml of Barium chloride solution is added. 2 Silver Nitrate Test: To a few drops of the extract dilute Nitric Acid is added until the effervescence ceases and 2 ml of Silver Nitrate solution is added. 3 Lead Acetate Test To a few drops of the extract dilute acetic acid is added until the effervescence ceases and 2 ml of Lead acetate solution is added 4 Brown Ring Test: To a few drops of extract dilute Sulphuric acid is added until the effervescence ceases, then freshly prepared FeSO4 is added and then conctrated Sulphuric acid is added drop by drop along the sides of the test tube. IDENTIFICATION OF THE BASIC RADICALS PREPARATION OF ORIGINAL SOLUTION: The orginal solution prepared by dissolving a small amound of salt in water ZERO GROUP 1 To a few drops of the original solution Sodium Hydroxide and Nessler's reagent are added. No curdy white precipitate, No curdy white precipitate, Absence of Chloride. Absence of Sulphate Absence of Nitrate. Absence of	minute			
To a few drops of the extract dilute Nitric	1	To a few drops of the extract, dilute Hydrochloric Acid is added until the effervescence ceases and 2 ml of Barium	No white precipitate.	·
Absence of sulphate To a few drops of the extract dilute acetic acid is added until the effervescence ceases and 2 ml of Lead acetate solution is added	2	To a few drops of the extract dilute Nitric Acid is added until the effervescence ceases	No curdy white precipitate,	Absence of Chloride .
4 Brown Ring Test: To a few drops of extract dilute Sulphuric acid is added until the effervescence ceases, then freshly prepared FeSO4 is added and then conctrated Sulphuric acid is added drop by drop along the sides of the test tube. IDENTIFICATION OF THE BASIC RADICALS PREPARATION OF ORIGINAL SOLUTION: The orginal solution prepared by dissolving a small amound of salt in water ZERO GROUP 1 To a few drops of the original solution Sodium Hydroxide and Nessler's reagent are added. Reddish brown precipitate. Ammonium is confirmed.	3	Lead Acetate Test To a few drops of the extract dilute acetic acid is added until the effervescence ceases and 2	No precipitate	Absence of sulphate
IDENTIFICATION OF THE BASIC RADICALS PREPARATION OF ORIGINAL SOLUTION: The orginal solution prepared by dissolving a small amound of salt in water ZERO GROUP 1 To a few drops of the original solution Sodium Hydroxide and Nessler's reagent are added. Reddish brown precipitate. Ammonium is confirmed.	4	Brown Ring Test: To a few drops of extract dilute Sulphuric acid is added until the effervescence ceases, then freshly prepared FeSO ₄ is added and then conctrated Sulphuric acid is added	No brown ring.	Absence of Nitrate.
To a few drops of the original solution Sodium Hydroxide and Nessler's reagent are added. ZERO GROUP Reddish brownprecipitate. Reddish brownprecipitate. Ammonium is confirmed.		IDENTIFICATION OF ORIGINAL SOLUTION:		
To a few drops of the original solution Sodium Hydroxide and Nessler's reagent are added. Reddish brownprecipitate. Reddish brownprecipitate. Ammonium is confirmed.	1110 01	Burn ooranon brobarea by aissorring a sman		
	1	solution Sodium Hydroxide and Nessler's		Ammonium is confirmed.
			GROUP SEPARATION	

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2	To a few drops of the original solution 2 ml of dilute HCl is added	No characteristic precipitate.	Absence of First Group (Lead).
3	To a few drops of the original solution 1 ml NH ₄ Cl and 2 ml NH ₄ OH solutions are added.	No characteristic precipitate.	Absence of Third Group (Aluminium).
4	To a few drops of the original solution 1 ml NH ₄ Cl and 2 ml NH ₄ OH solutions are added and H ₂ S gas is passed.	No characteristic precipitate.	Absence of Fourth Group (Zinc).
5	To a few drops of the original solution 1 ml NH ₄ Cl, 2 ml NH ₄ OH and 2 ml (NH ₄) ₂ CO ₃ solutions are added.	No characteristic precipitate	Absence of Fifth Group (Calcium and Barium).
6	To a few drops of the original solution 1 ml NH ₄ Cl, 2 ml NH ₄ OH and 2 ml Di Sodium Hydrogen Phosphate are added.	No characteristic precipitate.	Absence of Sixth Group (Magnesium).
	CONFIRMAT	ORY TESTS FOR BASIC RADICALS	
1	To a few drops of the original solution Sodium Hydroxide and Nessler's reagent are added.	Reddish brown precipitate.	Ammonium is confirmed.

RESULT

The given simple salt contains **1.** Basic Radical : **Ammonium** 2. Acid Radical : Carbonate

The given simple salt is : $\mbox{\bf Ammonium Carbonate}$

Simple salt analysis –VII PRELIMINARY TESTS DATE

S. No.	EXPERIMENT	OBSERVATION	INFERENCE
1	a)Colour Colour of the salt is noted	Colourless	Absence of copper and iron salt
	b)Appearance Appearance of the salt is noted	crystaline	May be Sulphate, Nitrate, Chloride
2	Solubility A little of the salt is shaken with water.	Soluble	May be Sulphate, Nitrate, Chloride or Ammonium Carbonate.
3	Action Of Heat: A small amount of the salt is heated gently in a dry test tube.	No characteristic change.	Absence of Carbonate, Nitrate, Ammonium and Zinc.
4	Flame Test: A small amount of the salt is made into a paste with conc. HCl in a watch glass and introduced into the non-luminous part of the Bunsen flame.	No characteristic coloured flame	Absence of Calcium and Barium.
5	Ash Test: A filter paper is soaked into a paste of the salt with conc. HCl / HNO ₃ And Cobalt Nitrate solution in a watch glass and burnt.	Pink ash.	Presence of Magnesium .
		ICATION OF ACID RADICALS	1
6	Action Of Dil Hcl To a small amount of dilute HCl the salt is added.	No characteristic change.	Absence of Sulphide and Carbonate.
7	Copper Turnings Test: A small amount of the salt is heated with Copper Turnings and a few drops of conc. Sulphuric acid.	No reddish brown gas is evolved.	Absence of Nitrate.
8	Action Of NaOH: A small amount of the salt is heated with Sodium Hydroxide.	No pungent smelling gas.	Absence of Ammonium.
9	Chromyl Chloride Test: To a small amount of the salt a pinch of Potassium Dichromate is added and	No Red orange vapours.	Absence of chloride.

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	heated with few drops of conc. Sulphuric		
	acid.		
	ARATION OF SODIUM CARBONATE EXTRACT:		
	ll amount of salt is mixed with twice the amount		lled water is added, boiled for 10
minut	es, cooled and filtered. The filtrate is called "SOD		
1	Barium Chloride Test:	A white precipitate, insoluble in	Sulphate is confirmed.
	To a few drops of the extract, dilute	conc. HCl.	
	Hydrochloric Acid is added until the		
	effervescence ceases and 2 ml of Barium		
	chloride solution is added.		
2	Silver Nitrate Test:	No precipitate.	Absence of chloride/ sulphide
	To a few drops of the extract dilute Nitric		
	Acid is added until the effervescence ceases		
	and 2 ml of Silver Nitrate solution is added.		
3	Lead Acetate Test	White precipitate soluble in	Presence of sulphate
	To a few drops of the extract dilute acetic	ammonium acetate and sodium	
	acid is added until the effervescence ceases	hydroxide is obtained	
	and 2 ml of Lead acetate solution is added		
4	Brown Ring Test:	No brown ring.	Absence of Nitrate.
	To a few drops of extract dilute Sulphuric		
	acid is added until the effervescence ceases,		
	then freshly prepared FeSO ₄ is added and		
	then conctrated Sulphuric acid is added		
	drop by drop along the sides of the test tube.		
	ARATION OF ORIGINAL SOLUTION: rginal solution prepared by dissolving a smal	l amound of salt in water ZERO GROUP	
1	T =	ZLIO UKOUI	
1	To a few drone of the original	No Raddish brown precipitate	Absence of Ammonium
	To a few drops of the original	No Reddish brown precipitate.	Absence of Ammonium
	solution Sodium Hydroxide and Nessler's	No Reddish brown precipitate.	Absence of Ammonium
	solution Sodium Hydroxide and Nessler's reagent are added.		Absence of Ammonium
2	solution Sodium Hydroxide and Nessler's reagent are added.	GROUP SEPARATION	
2	solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml		Absence of Ammonium Absence of First Group (Lead).
	solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml of dilute HCl is added	GROUP SEPARATION No characteristic precipitate.	Absence of First Group (Lead).
2	solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml	GROUP SEPARATION	Absence of First Group (Lead). Absence of Third Group
3	solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added.	GROUP SEPARATION No characteristic precipitate. No characteristic precipitate.	Absence of First Group (Lead). Absence of Third Group (Aluminium).
	solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml	GROUP SEPARATION No characteristic precipitate.	Absence of First Group (Lead). Absence of Third Group
3	solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH ₄ Cl and 2 ml NH ₄ OHsolutions are added. To a few drops of the original solution 1 ml NH ₄ Cl and 2 ml NH ₄ OHsolutions are added	GROUP SEPARATION No characteristic precipitate. No characteristic precipitate.	Absence of First Group (Lead). Absence of Third Group (Aluminium).
3	solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH ₄ Cl and 2 ml NH ₄ OHsolutions are added. To a few drops of the original solution 1 ml NH ₄ Cl and 2 ml NH ₄ OHsolutions are added and H ₂ S gas is passed.	GROUP SEPARATION No characteristic precipitate. No characteristic precipitate. No characteristic precipitate.	Absence of First Group (Lead). Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc).
3	solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added and H2S gas is passed. To a few drops of the original solution 1 ml	GROUP SEPARATION No characteristic precipitate. No characteristic precipitate.	Absence of First Group (Lead). Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc).
3 4	solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added and H2S gas is passed. To a few drops of the original solution 1 ml NH4Cl, 2 ml NH4OH and 2 ml (NH4)2CO3	GROUP SEPARATION No characteristic precipitate. No characteristic precipitate. No characteristic precipitate.	Absence of First Group (Lead). Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc).
3 4 5	solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added and H2S gas is passed. To a few drops of the original solution 1 ml NH4Cl, 2 ml NH4OH and 2 ml (NH4)2CO3 solutions are added.	GROUP SEPARATION No characteristic precipitate. No characteristic precipitate. No characteristic precipitate. No characteristic precipitate.	Absence of First Group (Lead). Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc). Absence of Fifth Group (Calcium and Barium).
3 4	solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added and H2S gas is passed. To a few drops of the original solution 1 ml NH4Cl, 2 ml NH4OH and 2 ml (NH4)2CO3 solutions are added. To a few drops of the original solution 1 ml	GROUP SEPARATION No characteristic precipitate. No characteristic precipitate. No characteristic precipitate.	Absence of First Group (Lead). Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc). Absence of Fifth Group (Calcium and Barium). Presence of Sixth Group
3 4 5	solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added and H2S gas is passed. To a few drops of the original solution 1 ml NH4Cl, 2 ml NH4OH and 2 ml (NH4)2CO3 solutions are added. To a few drops of the original solution 1 ml NH4Cl, 2 ml NH4OH and 2 ml Di Sodium	GROUP SEPARATION No characteristic precipitate. No characteristic precipitate. No characteristic precipitate. No characteristic precipitate.	Absence of First Group (Lead). Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc). Absence of Fifth Group (Calcium and Barium).
3 4 5	solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added and H2S gas is passed. To a few drops of the original solution 1 ml NH4Cl, 2 ml NH4OH and 2 ml (NH4)2CO3 solutions are added. To a few drops of the original solution 1 ml NH4Cl, 2 ml NH4OH and 2 ml Di Sodium Hydrogen Phosphate are added.	GROUP SEPARATION No characteristic precipitate. No characteristic precipitate. No characteristic precipitate. No characteristic precipitate. White precipitate.	Absence of First Group (Lead). Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc). Absence of Fifth Group (Calcium and Barium). Presence of Sixth Group
3 4 5 6	solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added and H2S gas is passed. To a few drops of the original solution 1 ml NH4Cl, 2 ml NH4OH and 2 ml (NH4)2CO3 solutions are added. To a few drops of the original solution 1 ml NH4Cl, 2 ml NH4OH and 2 ml Di Sodium Hydrogen Phosphate are added.	GROUP SEPARATION No characteristic precipitate. No characteristic precipitate. No characteristic precipitate. No characteristic precipitate. White precipitate. ORY TESTS FOR BASIC RADICALS	Absence of First Group (Lead). Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc). Absence of Fifth Group (Calcium and Barium). Presence of Sixth Group (Magnesium).
3 4 5	solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added and H2S gas is passed. To a few drops of the original solution 1 ml NH4Cl, 2 ml NH4OH and 2 ml (NH4)2CO3 solutions are added. To a few drops of the original solution 1 ml NH4Cl, 2 ml NH4OH and 2 ml Di Sodium Hydrogen Phosphate are added. CONFIRMAT	GROUP SEPARATION No characteristic precipitate. No characteristic precipitate. No characteristic precipitate. No characteristic precipitate. White precipitate.	Absence of First Group (Lead). Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc). Absence of Fifth Group (Calcium and Barium). Presence of Sixth Group
3 4 5 6	solution Sodium Hydroxide and Nessler's reagent are added. To a few drops of the original solution 2 ml of dilute HCl is added To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added. To a few drops of the original solution 1 ml NH4Cl and 2 ml NH4OHsolutions are added and H2S gas is passed. To a few drops of the original solution 1 ml NH4Cl, 2 ml NH4OH and 2 ml (NH4)2CO3 solutions are added. To a few drops of the original solution 1 ml NH4Cl, 2 ml NH4OH and 2 ml Di Sodium Hydrogen Phosphate are added.	GROUP SEPARATION No characteristic precipitate. No characteristic precipitate. No characteristic precipitate. No characteristic precipitate. White precipitate. ORY TESTS FOR BASIC RADICALS	Absence of First Group (Lead). Absence of Third Group (Aluminium). Absence of Fourth Group (Zinc). Absence of Fifth Group (Calcium and Barium). Presence of Sixth Group (Magnesium).

RESULT

The given simple salt contains

1. Basic Radical : Magnesium

2. Acid Radical : Sulphate

The given simple salt is : Magnesium Sulphate

Simple salt analysis -VIII

PRELIMINARY TESTS DATE S. No. EXPERIMENT OBSERVATION INFERENCE 1 a)Colour Absence of copper and iron salt Colourless Colour of the salt is noted b)Appearance Appearance of the salt is noted crystaline May be Sulphate, Nitrate, Chloride May be Sulphate, Nitrate, Chloride or Solubility Soluble 2

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	A little of the salt is shaken with water.		Ammonium Carbonate.
3	Action Of Heat:	No characteristic change.	Absence of Carbonate, Nitrate,
	A small amount of the salt is heated gently		Ammonium and Zinc.
	in a dry test tube.		
4	Flame Test:	Grassy green flame.	Presence of Barium .
•	A small amount of the salt is made into a	- 229 6 22 2000	
	paste with conc. HCl in a watch glass and		
	introduced into the non-luminous part of the		
	Bunsen flame.		1
5	Ash Test:	No characteristic coloured ash.	Absence of Zinc, Aluminium and
	A filter paper is soaked into a paste of the		Magnesium.
	salt with conc. HCl / HNO ₃		
	And Cobalt Nitrate solution in a watch glass		
	and burnt.		
		TICATION OF ACID RADICALS	
6	Action Of Dil Hcl	No characteristic change.	Absence of Sulphide and Carbonate.
U	To a small amount of dilute HCl the salt is	1.0 characteristic change.	Tibbeliee of bulpfilde and carbonate.
	added.	11111	11
7	Copper Turnings Test:	No reddish brown gas is evolved.	Absence of Nitrate.
	A small amount of the salt is heated		
	with Copper Turnings and a few drops		
	of conc. Sulphuric acid.		
8	Action Of NaOH:	No pungent smelling gas.	Absence of Ammonium.
J	A small amount of the salt is heated		
	with Sodium Hydroxide.	D. J	Chlorida C
9	Chromyl Chloride Test:	Red orange vapours evolved are	Chloride is confirmed.
	To a small amount of the salt a pinch	passed through water to get a	
	of Potassium Dichromate is added and	yellow solution, which on	I P
	heated with few drops of conc. Sulphuric	adding Lead acetate forms	
	acid.	a yellowprecipitate	
PREPA	RATION OF SODIUM CARBONATE EXTRACT:		·
	l amount of salt is mixed with twice the amount of	of sodium carbonate and 20ml of dist	illed water is added, boiled for 10
	es, cooled and filtered. The filtrate is called "SOD "		
	Barium Chloride Test:		Absongs of Cylinhata
1		No white precipitate.	Absence of Sulphate
	To a few drops of the extract, dilute		
	Hydrochloric Acid is added until the		
	effervescence ceases and 2 ml of Barium		
	chloride solution is added.		
2	Silver Nitrate Test:	A curdy white precipitate,	Presence of Chloride .
	To a few drops of the extract dilute Nitric	soluble in excess of Ammonium	
	Acid is added until the effervescence ceases	hydroxide.	
	and 2 ml of Silver Nitrate solution is added.	ily al oxide.	
2		No procipitate	Abaanaa of sululusta
3	Lead Acetate Test	No precipitate	Absence of sulphate
	To a few drops of the extract dilute acetic acid		
	is added until the effervescence ceases and 2		
	ml of Lead acetate solution is added		
4	Brown Ring Test:	No brown ring.	Absence of Nitrate.
	To a few drops of extract dilute Sulphuric		
	acid is added until the effervescence ceases,		
	then freshly prepared FeSO ₄ is added and		
	then conctrated Sulphuric acid is added		
	drop by drop along the sides of the test tube.		
	IDENTIFICA	ATION OF THE BASIC RADICALS	
PREPA	ARATION OF ORIGINAL SOLUTION:		
	ginal solution prepared by dissolving a small	amound of salt in water	
		ZERO GROUP	
1	To a few drops of the original	No Reddish brown precipitate.	Absence of Ammonium
T		No Reduish brown precipitate.	Absence of Allinomuli
	solution Sodium Hydroxide and Nessler's		
	reagent are added.	CROVED CERTAIN ATTACK	L
		GROUP SEPARATION	
2	To a few drops of the original solution 2 ml	No characteristic precipitate.	Absence of First Group (Lead).
	of dilute HCl is added		
3	To a few drops of the original solution 1 ml	No characteristic precipitate.	Absence of Third Group
5	NH ₄ Cl and 2 ml NH ₄ OH solutions are added.	characteristic precipitate.	(Aluminium).
1		No characteristic was similate	
4	To a few drops of the original solution 1 ml	No characteristic precipitate.	Absence of Fourth Group (Zinc).
	NH ₄ Cl and 2 ml NH ₄ OH solutions are added		
	and H₂S gas is passed.		

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	NH ₄ Cl, 2 ml NH ₄ OH and 2 ml (NH ₄) ₂ CO ₃ solutions are added.		Group (Calcium orBarium).
	CONFIRMAT	ORY TESTS FOR BASIC RADICALS	
1	To a few drops of the original solution Potassium Chromate is added.	Yellow precipitate, soluble in acid.	Barium is confirmed.

RESULT

The given simple salt contains

1. Basic Radical : Barium 2. Acid Radical : Chloride

The given simple salt is: Barium Chloride

	DATE PRELIMIN	Simple salt analysis -IX IARY TESTS	
S. No.	EXPERIMENT	OBSERVATION	INFERENCE
1	a)Colour Colour of the salt is noted	Colourless	Absence of copper and iron salt
	b)Appearance Appearance of the salt is noted	crystaline	May be Sulphate, Nitrate, Chloride
2	Solubility A little of the salt is shaken with water.	Soluble	May be Sulphate, Nitrate, Chloride or Ammonium Carbonate.
3	Action Of Heat: A small amount of the salt is heated gently in a dry test tube.	No characteristic change.	Absence of Carbonate, Nitrate, Ammonium and Zinc.
4	Flame Test: A small amount of the salt is made into a paste with conc. HCl in a watch glass and introduced into the non-luminous part of the Bunsen flame.	No characteristic coloured flame	Absence of Calcium and Barium.
5	Ash Test: A filter paper is soaked into a paste of the salt with conc. HCl / HNO ₃ And Cobalt Nitrate solution in a watch glass and burnt.	B <u>l</u> ue ash.	Presence of A<u>I</u>uminium .
		TICATION OF ACID RADICALS	
6	Action Of Dil Hcl To a small amount of dilute HCl the salt is added.	No characteristic change.	Absence of Sulphide and Carbonate.
7	Copper Turnings Test: A small amount of the salt is heated with Copper Turnings and a few drops of conc. Sulphuric acid.	No reddish brown gas is evolved.	Absence of Nitrate.
8	Action Of NaOH: A small amount of the salt is heated with Sodium Hydroxide.	No pungent smelling gas.	Absence of Ammonium.
9	Chromyl Chloride Test: To a small amount of the salt a pinch of Potassium Dichromate is added and heated with few drops of conc. Sulphuric acid.	No Red orange vapours.	Absence of chloride.
	RATION OF SODIUM CARBONATE EXTRACT:		
	l amount of salt is mixed with twice the amount		lled water is added, boiled for 10
	es, cooled and filtered. The filtrate is called "SOD "		
1	Barium Chloride Test: To a few drops of the extract, dilute Hydrochloric Acid is added until the effervescence ceases and 2 ml of Barium chloride solution is added.	A white precipitate, insoluble in conc. HCl.	Sulphate is confirmed.

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2	Silver Nitrate Test:	No precipitate.	Absence of chloride/ sulphide
2		No precipitate.	Absence of chioride/ surplinde
	To a few drops of the extract dilute Nitric Acid is added until the effervescence ceases		
2	and 2 ml of Silver Nitrate solution is added.	TATILLE CONTROL OF THE CONTROL OF TH	D
3	Lead Acetate Test	White precipitate soluble in	Presence of sulphate
	To a few drops of the extract dilute acetic	ammonium acetate and sodium	
	acid is added until the effervescence ceases	hydroxide is obtained	
	and 2 ml of Lead acetate solution is added		
4	Brown Ring Test:	No brown ring.	Absence of Nitrate.
	To a few drops of extract dilute Sulphuric		
	acid is added until the effervescence ceases,		
	then freshly prepared FeSO 4 is added and		
	then conctrated Sulphuric acid is added		
	drop by drop along the sides of the test tube.		
	IDENTIFICA	ATION OF THE BASIC RADICALS	
	ARATION OF ORIGINAL SOLUTION:		
The or	ginal solution prepared by dissolving a small	amound of salt in water	
		ZERO GROUP	
1	To a few drops of the original	No Reddish brown precipitate.	Absence of Ammonium
	solution Sodium Hydroxide and Nessler's		
	reagent are added.		
		GROUP SEPARATION	
2	To a few drops of the original solution 2 ml	No characteristic precipitate.	Absence of First Group (Lead).
	of dilute HCl is added		
3	To a few drops of the original solution 1 ml	No characteristic precipitate.	Absence of Third Group
	1 0	1 1	-
	NH ₄ Cl and 2 ml NH ₄ OH solutions are added.		(Aluminium).
	NH4Cl and 2 ml NH4OHsolutions are added. CONFIRMAT	ORY TESTS FOR BASIC RADICALS	(Aluminium).
1	CONFIRMAT		
1	CONFIRMAT To a few drops of the original	ORY TESTS FOR BASIC RADICALS A bright red lake.	(Aluminium). Aluminium is confirmed.
1	CONFIRMAT		

RESULT

added.

The given simple salt contains

1. Basic Radical : Aluminium

2. Acid Radical : Sulphate

The given simple salt is : Aluminium Sulphate

EXPERIMENT

Simple salt analysis -X PRELIMINARY TESTS

OBSERVATION

INFERENCE

|--|

1	a)Colour Colour of the salt is noted	Colourless	Absence of copper and iron salt
	b)Appearance Appearance of the salt is noted	powdery	May be carbonate and sulphide
2	Solubility A little of the salt is shaken with water.	insoluble	May be carbonate and sulphide
3	Action Of Heat: A small amount of the salt is heated gently in a dry test tube.	No characteristic change.	Absence of Carbonate, Nitrate, Ammonium and Zinc.
4	Flame Test: A small amount of the salt is made into a paste with conc. HCl in a watch glass and introduced into the non-luminous part of the Bunsen flame.	No characteristic coloured flame	Absence of Calcium and Barium.
5	Ash Test: A filter paper is soaked into a paste of the salt with conc. HCl / HNO ₃ And Cobalt Nitrate solution in a watch glass and burnt.	Gree <u>n</u> ash.	Presence of Zi<u>n</u>c.

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		ICATION OF ACID RADICALS	
6	Action Of Dil Hcl To a small amount of dilute HCl the salt is added.	Rotten egg smelling gas turning Lead acetate paper black.	Sulphide is confirmed.
7	Copper Turnings Test: A small amount of the salt is heated with Copper Turnings and a few drops of conc. Sulphuric acid.	No reddish brown gas is evolved.	Absence of Nitrate.
8	Action Of NaOH: A small amount of the salt is heated with Sodium Hydroxide.		
9	Chromyl Chloride Test: To a small amount of the salt a pinch of Potassium Dichromate is added and heated with few drops of conc. Sulphuric acid.	No Red orange vapours.	Absence of chloride.
PREPA	RATION OF SODIUM CARBONATE EXTRACT:		
	l amount of salt is mixed with twice the amount or soled and filtered. The filtrate is called "SOD I"		lled water is added, boiled for 10
1	Barium Chloride Test: To a few drops of the extract, dilute Hydrochloric Acid is added until the effervescence ceases and 2 ml of Barium chloride solution is added.	No white precipitate.	Absence of Sulphate
2	Silver Nitrate Test: To a few drops of the extract dilute Nitric Acid is added until the effervescence ceases and 2 ml of Silver Nitrate solution is added.	A black precipitate.	Presence of Sulphide .
3	Lead Acetate Test To a few drops of the extract dilute acetic acid is added until the effervescence ceases and 2 ml of Lead acetate solution is added	A black precipitate insoluble in hot nitric acid is obtained	Presence of sulphide
4	Brown Ring Test: To a few drops of extract dilute Sulphuric acid is added until the effervescence ceases, then freshly prepared FeSO ₄ is added and then conctrated Sulphuric acid is added	No brown ring.	Absence of Nitrate.
	drop by drop along the sides of the test tube.	ATTION OF THE PACIC PARICAL C	
	ARATION OF ORIGINAL SOLUTION:	ATION OF THE BASIC RADICALS	
ine or	ginal solution prepared by dissolving a small	ZERO GROUP	
1	To a few drops of the original solution Sodium Hydroxide and Nessler's reagent are added.	No Reddish brown precipitate.	Absence of Ammonium
		GROUP SEPARATION	
2	To a few drops of the original solution 2 ml of dilute HCl is added	No characteristic precipitate.	Absence of First Group (Lead).
3	To a few drops of the original solution 1 ml NH ₄ Cl and 2 ml NH ₄ OH solutions are added.	No characteristic precipitate.	Absence of Third Group (Aluminium).
4	To a few drops of the original solution 1 ml NH ₄ Cl and 2 ml NH ₄ OHsolutions are added and H ₂ S gas is passed.	Dirty whiteprecipitate.	Presence of Fourth Group (Zinc).
		ORY TESTS FOR BASIC RADICALS	
1	To a few drops of the original solution Potassium Ferro cyanide is added.	White precipitatesoluble in excess of Sodium hydroxide &insoluble in dilute acid.	Zinc is confirmed.
	RESULT	aciu.	<u> </u>

RESULT

The given simple salt contains

1. Basic Radical : Zinc

2. Acid Radical : Sulphide

The given simple salt is : Zinc Sulphide

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